

CIRCLE YOUR COURSE SECTION IN THE LIST BELOW

202-12 W 10-12:50

202-13 W 2-2:50

202-14 W 3-3:50

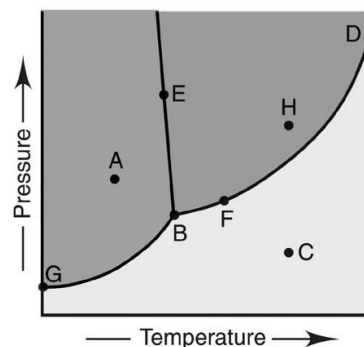
Section 1. Phases and Phase Changes (14 pts)

1. Identify the following on the phase diagram to the right:

A. Phase at point H

B. Term for point B

C. Phase change upon lowering P from point A



2. Given an example of a phase change that is exothermic.

3. If a substance has $\Delta H^\circ_{\text{vap}}$ of 30.72 kJ/mol and $\Delta H^\circ_{\text{fus}}$ = 9.87 kJ/mol, what is $\Delta H^\circ_{\text{cond}}$?

4. The boiling point of ethane is -88.6 °C. The boiling point of ethanol is 78.3 °C. Which compound has the higher vapor pressure at a given temperature?

Section 2. Intermolecular Forces (18 pts)

5. List the intermolecular forces operating in the following; if more than one is operating circle the strongest.

A. BCl_3

B. NH_3

6. Identify the primary intermolecular forces (IF) operating in the following solutions. Each line is one IF.

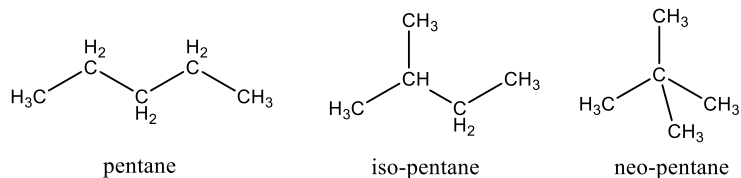
A. CCl_4 in H_2O

B. KNO_3 in H_2O

7. Rank the following gases from the lowest to highest solubility in water? CO , SO_2 , N_2

Section 3. Properties of solids and liquids (14 pts)

8. Rank the following liquids from the lowest to highest surface tension? pentane, iso-pentane, neo-pentane



9. Solid iron can form either a body-centered cubic lattice or face-centered cubic lattice.

How many iron atoms are in one unit cell of each type?

Body-centered cubic lattice _____ face-centered cubic lattice _____

10. Which cubic lattice has a coordination number of 6?

Section 4. Mixtures and solutions (14 pts)

11. List two ways to increase the solubility of O_2 gas in water.

12. Rank the solutions from lowest to highest vapor pressure. 0.1 m sucrose ($C_6H_{12}O_6$), 0.1 m KBr, 0.1 m $CaCl_2$

13. Compound AX is an ionic compound that dissolves in water to yield $A^+_{(aq)}$ and $X^-_{(aq)}$.
Write a chemical equation for ΔH°_{hydr} of AX.

Section 5. Calculations (24 pts). YOU MUST SHOW YOUR WORK FOR CREDIT.

14. How much heat (in kJ) is absorbed when 5.28 grams of bromine (Br_2) evaporates at atmospheric pressure?

Standard enthalpies for Br_2 (in kJ/mol)	
ΔH°_{vap}	29.6
ΔH°_{fus}	11.6
ΔH°_{sub}	41.2

15. At 25 °C and 785 Torr, the solubility of carbon dioxide in water is 0.0351 M.
What is its solubility at 25 °C and 1390 Torr?

16. Using the data to the right, calculate the relative humidity when the atmospheric temperature is 20 °C and the dewpoint is 10 °C.

Temperature (°C)	Equilibrium Vapor Pressure (Torr)
10	9.2
15	12.8
20	17.5
25	23.8

17. Calculate the concentration of a 7.82×10^{-4} M aqueous solution of fluoride (F^-) in parts per million. Assume the density of the solution is 1.00 g/mL.

Section 6. Essay (16 pts). Answer in paragraph form comprised of 3-5 full sentences.

18. A student drops a raisin into a beaker containing distilled (pure) water.

After 1 hour, the raisin has increased in size. Explain.

19. A student uses Buchner funnel with a Buchner flask connected to a vacuum aspirator to filter a methanol (CH_3OH) suspension containing solid product. The student leaves the apparatus connected to the vacuum for 15 minutes after the filtration to allow the solid sample to dry. During this time, the student notices a colorless liquid on the outside of the Buchner flask. Explain.

THIS PAGE CAN BE REMOVED FROM THE EXAM

USE THE BACK FOR SCRATCH PAPER IF NEEDED

The Periodic Table of the Elements

1 H Hydrogen 1.00794																	2 He Helium 4.003										
3 Li Lithium 6.941	4 Be Beryllium 9.012182											5 B Boron 10.811	6 C Carbon 12.0107	7 N Nitrogen 14.00674	8 O Oxygen 15.9994	9 F Fluorine 18.9984032	10 Ne Neon 20.1797										
11 Na Sodium 22.989770	12 Mg Magnesium 24.3050											13 Al Aluminum 26.981538	14 Si Silicon 28.0855	15 P Phosphorus 30.973761	16 S Sulfur 32.066	17 Cl Chlorine 35.4527	18 Ar Argon 39.948										
19 K Potassium 39.0983	20 Ca Calcium 40.078	21 Sc Scandium 44.955910	22 Ti Titanium 47.867	23 V Vanadium 50.9415	24 Cr Chromium 51.9961	25 Mn Manganese 54.938049	26 Fe Iron 55.845	27 Co Cobalt 58.933200	28 Ni Nickel 58.6934	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.723	32 Ge Germanium 72.61	33 As Arsenic 74.92160	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80										
37 Rb Rubidium 85.4678	38 Sr Strontium 87.62	39 Y Yttrium 88.90585	40 Zr Zirconium 91.224	41 Nb Niobium 92.90638	42 Mo Molybdenum 95.94	43 Tc Technetium (98)	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.90550	46 Pd Palladium 106.42	47 Ag Silver 107.8682	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.60	53 I Iodine 126.90447	54 Xe Xenon 131.29										
55 Cs Cesium 132.90545	56 Ba Barium 137.327	57 La Lanthanum 138.9055	72 Hf Hafnium 178.49	73 Ta Tantalum 180.9479	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.078	79 Au Gold 196.96655	80 Hg Mercury 200.59	81 Tl Thallium 204.3833	82 Pb Lead 207.2	83 Bi Bismuth 208.98038	84 Po Polonium (209)	85 At Astatine (210)	86 Rn Radon (222)										
87 Fr Francium (223)	88 Ra Radium (226)	89 Ac Actinium (227)	104 Rf Rutherfordium (261)	105 Db Dubnium (262)	106 Sg Seaborgium (263)	107 Bh Bohrium (262)	108 Hs Hassium (265)	109 Mt Meitnerium (266)	110 (269)	111 (272)	112 (277)	113	114														
58 Ce Cerium 140.116	59 Pr Praseodymium 140.90765	60 Nd Neodymium 144.24	61 Pm Promethium (145)	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.92534	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93032	68 Er Erbium 167.26	69 Tm Thulium 168.93421	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967	90 Th Thorium 232.0381	91 Pa Protactinium 231.03588	92 U Uranium 238.0289	93 Np Neptunium (237)	94 Pu Plutonium (244)	95 Am Americium (243)	96 Cm Curium (247)	97 Bk Berkelium (247)	98 Cf Californium (251)	99 Es Einsteinium (252)	100 Fm Fermium (257)	101 Md Mendelevium (258)	102 No Nobelium (259)	103 Lr Lawrencium (262)